

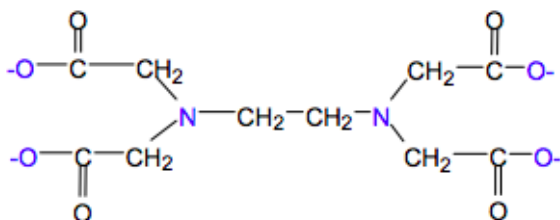
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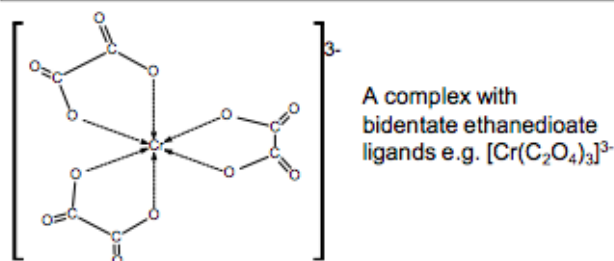
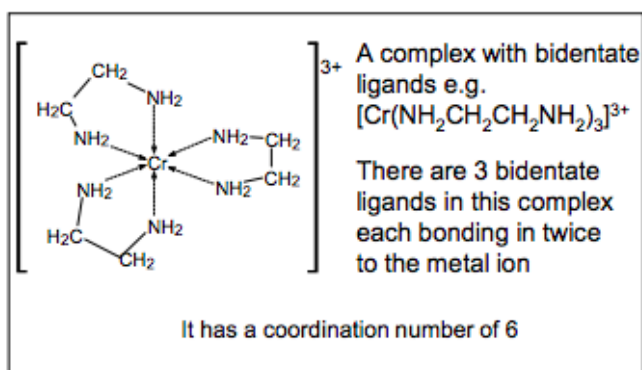
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ligands can be **unidentate** (e.g. H_2O , NH_3 and Cl^-) which can form one coordinate bond per ligand or **bidentate** (e.g. $\text{NH}_2\text{CH}_2\text{CH}_2\text{NH}_2$ and ethanedioate ion $\text{C}_2\text{O}_4^{2-}$) which have two atoms with lone pairs and can form two coordinate bonds per ligand or **multidentate** (e.g. EDTA^{4-} which can form six coordinate bonds per ligand).

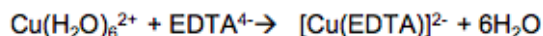
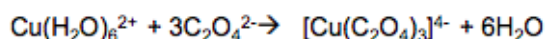
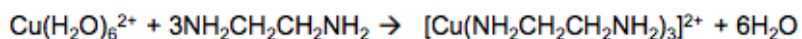
The EDTA^{4-} anion has the formula



with six donor sites (4O and 2N) and forms a 1:1 complex with metal(II) ions



Equations to show formation of bidentate and multidentate complexes



Learn the two bidentate ligands mentioned above but it is not necessary to remember the structure of EDTA

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